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Episode 6: Cartridge Pressure Testing | Research \u0026amp; Development **Solution**

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Basic concepts and definitions - Part 1

The Energy Equation | First Law of Thermodynamics | Work | Heat Transfer | Internal Energy | Enthalpy

Calculating k_{see} (ideal gas or ideal solution) - Chemical Engineering Thermodynamics **Lec 04 Energy, Work and Heat, Point and Path functions CH6503**
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 get $m = \rho V = 1 \text{ kg/m}^3 \times$
 $0.001 \text{ m}^3 = 0.001 \text{ kg}$
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 Table A.5 is $\rho = 1.17$
 kg/m^3 at 100 kPa, 25°C.
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 $V = 110 \text{ kg/m}^3 \times 100 \text{ 000}$

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 $= \text{Constant} = v_0$. If you
 include that v increases a
 little with T then: $v = v_0 +$
 $C (T - T_0)$ where C is a
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 $\text{kg/m}^3 \times 100\,000 \text{ m}^3 =$
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 this in perspective a
 power plant that
 generates 2000 MW by
 burning coal would make
 about 20 million tons of
 carbon-dioxide a year.

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 above mass so it is nearly
 impossible to store all the
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