

# Chapter 122 Stoichiometric Calculations

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atoms in 0.5 mole of oxalic acid =  $6.022 \times 10^{23} \times 0.5 \times 2 = 6.022 \times 10^{23}$ . 4. 0.5216g of a solid mixture containing  $\text{Na}_2\text{SO}_4$  is dissolved in water and treated with an excess of  $\text{BaCl}_2$ , resulting in the precipitation of 0.6168g of  $\text{BaSO}_4$ . What is Stoichiometry? Balancing Equations, Stoichiometric ... One can find the percentage of the mass of a compound that comes from each of the elements in the compound by using this equation: % element = (number of atoms)(atomic weight) (FW of the compound)  $\times 100$ . Stoichiometry. Percent Composition. Chapter 3 Stoichiometry - Chemistry 12.2 Stoichiometric calculations. As stated before, stoichiometry represents the quantitative relationships between amount of reactants used and amount of products formed. Remember that all stoichiometric calculations must begin with.... Balancing! An incorrectly balanced equation is going to give you incorrect predictions of products

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Stoichiometric Calculations Summary Stoichiometric Calculations. Page 1 Page 2 Page 3 Applying Conversion Factors to Stoichiometry Now you're ready to use what you know about conversion factors to solve some stoichiometric problems in chemistry. Almost all stoichiometric problems can be solved in just four simple steps:

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One can find the percentage of the mass of a compound that comes from each of the elements in the compound by using this equation: % element = (number of atoms)(atomic weight) (FW of the compound)  $\times 100$ . Stoichiometry. Percent Composition.

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In chemical calculations, mole ratios are used to convert between moles of reactant and moles of

product, between moles of reactants, or between moles of products In a typical stoichiometric

problem, the given quantity is first converted to moles.

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Quantitative calculations that involve the stoichiometry of reactions in solution use volumes of

solutions of known concentration instead of masses of reactants or products. The coefficients in the

balanced chemical equation tell how many moles of reactants are needed and how many moles of

product can be produced.

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