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 1: Introduction and Definition of Terms †  
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 study of heat and work effects and  
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 $H \sum U \text{ ääääääää} \sum V LT = 0$  for an ideal gas  
 an  $H \sum U \text{ ääääääää} \sum T LV = n cv$  for an ideal  
 gas, the total differential for internal  
 energy for any change of state of an ideal  
 gas is  $dU = n cv dT$ .Gaskell -  
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 (David R. Gaskell) Problem 13.13: UC 2  
 can be equilibrated with UC and C at high  
 temperature and can be equilibrated with  
 U 2 C 3 and C at lower temperatures.

Calculate the maximum and minimum  
 temperatures at which UC 2 can  
 exist.Solved: Introduction To The  
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 Preliminaries † Settings Off @  
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 Needed for Problems ü Heat Capacities  
 The generic heat capacity  $C_p = a + bT$   
 $\text{ääääääää} 10^3 + c10^5$   
 $\text{ääääääääää} T^2$ ; The heat capacities  
 of various elements and compounds are  
 $C_p \text{Ags} = C_p \hat{e}. 8 \text{ a } \emptyset 21.30, b \emptyset 8.54, c \emptyset$   
 $1.51 <; C_p \text{Al} = C_p \hat{e}. 8 \text{ a } \emptyset 30.50, b \emptyset 0,$   
 $c \emptyset 0 <; C_p \text{Al} = C_p + 20.75T^2$  ...Solutions  
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Chapter 1: Introduction and Definition of Terms † History Thermodynamics began with the study of heat and work effects and relations between heat and work. Some early thermodynamics problems were for very practical problems. For example, in a steam engine heat is supplied to water to create steam. The steam is then used to turn an engine which does work.

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problem for the reaction  $\text{SO}_2 + 0.5 \text{O}_2 \rightleftharpoons \text{SO}_3$ . Lecture 16: 11.04.05 Single-Component phase diagrams ...

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Gaskell authored the textbooks Introduction to Metallurgical Thermodynamics, Introduction to the

Thermodynamics of Materials, and Introduction to Transport Phenomena in Materials Engineering. David E. Laughlin is the ALCOA Professor of Physical Metallurgy in the Department of Materials Science and Engineering of CMU and also has a courtesy appointment on the Electrical and Computer Engineering Department. *Introduction to the Thermodynamics of Materials (6th ed.)*

Gaskell. Chapter 2: The First Law of Thermodynamics † Ideal Gas Change of State ü Change in Internal Energy Because  $H \sum U \text{ åååååååå} \sum V LT = 0$  for an ideal gas an  $H \sum U \text{ åååååååå} \sum T LV = n cv$  for an ideal

gas, the total differential for internal energy for any change of state of an ideal gas is  $dU = n cv dT$ .

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