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 The responding variable was the color of the flame. The flame was orange at the end of each test because the wood splint started burning. The reason each of the elements produced a different color (or series of wavelengths) is because of the make up of the atom. ... Flame Test Lab Questions Answer Key: ...
 Flame Test Lab Questions Answer Key: This activity is called a flame test and it's a real procedure used in labs. Its purpose is to identify specific elements in a material. When the boric acid was in the flame, you probably notice a bright green portion of the flame. You may have seen it only briefly but it was there.
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 The objectives of this lab are to: Perform flame tests of metal cations in order to observe their characteristic colors, Perform calculations to determine the frequency and energy of the emitted ...
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 10. Hold the splint in the flame and record the color of the flame that is produced. 11. Using your data, identify the metal ion in your unknown solution. Flame Test Lab Activity Key Note: If chloride compounds are not available, metal nitrate compounds may be substituted. Use dilute or approximately 0.1 M solutions
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 Introduction to the Flame Test Lab: The Flame Test lab was an in-class lab where we tested chemicals in the flames to see the wide range of colors in the color spectrum. The secondary purpose of the lab was to identify unknown compounds that we would test and then guess as to what they were. The Flame Test lab was done in several parts.
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Flame Tests: A Favorite Chemistry Lab ... This flame test lab is always a favorite of mine, and a much loved lab by all of my students. The best time to use this lab is when teaching the following concepts: atomic structure, electron configurations, energy levels, ground state and excited state. ... [Chemistry Lab Flame Tests - Denton ISD](#)

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Flame Test Kit Student Laboratory Kit Introduction Just as a fingerprint is unique to each person, the color of light emitted by metals heated in a flame is unique to each metal. In this laboratory activity, the characteristic color of light emitted for calcium, copper, lithium, potassium, sodium, and strontium

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Amy Brown Science: Flame Tests: A Favorite Chemistry Lab

For this lab you must turn in the following items: Your data table recording flame test colors and wavelengths Answers to the following questions A one-page essay about how fireworks are made,

concentrating on how they produce different colors but including details of their basic construction and operation

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Atomic emission is utilized in everyday life in the laboratories of doctors and scientists. These individuals use the spectrum to test for the presence of elements such as sodium, potassium, iron, etc.. The color that the unknown solution emits will determine what element the

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