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$PV=nRT$ $P_1V_1/T_1 = P_2V_2/T_2$ 760 mm Hg = 1 atm $K = ^\circ C + 273$ 1000 mL = 1 L 22.4 L

= 1 mol (at STP) $R= 0.0821 \text{ atm}\cdot\text{L}$ $R=$

$8.314 \text{ kPa}\cdot\text{L}$ $R= 62.4 \text{ mm Hg}\cdot\text{L Mol}$

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