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Terms. Chapter 9: Chemical Names and
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PLAY. Monatomic Ion. Consists of a single
atom with a positive or negative charge
resulting from the loss of gain of one or
more valence electrons and behaves as a
unit, yet it still has a charge. (+ or -).
Determined by the amount of electrons
lost or gained. Chapter 9: Chemical Names
and Formulas Flashcards | Quizlet To name
a binary molecular compound, use the*

following guidelines: 1. Write the names of
the elements in the order listed in the
formula 2. Use prefixes appropriately to
indicate the number of each kind of atom
3. End the name of the second element
with the suffix -ide *Chemistry Chapter 9
Chemical Names and Formulas Flashcards
...First, confirm that the compound is a
binary molecular compound. Name the
elements in the order listed in the formula.
Use prefixes to indicate the number of
each kind of atom. Omit the prefix -mono
when the formula contains only one atom
of the first element in the name. The suffix
of the name of the second element is -
ide. Study Chapter 9: Chemical Names and
Formulas Flashcards ...1. when the name
of the anion (X) ends in -ide the acid name
begins with the prefix hydro-. The stem of
the anion has the suffix -ic and is followed
by the word acid (ex. HCl= hydrochloric
acid) 2. when the anion name ends in -ite
the acid name is the stem of the anion
with the suffix -ous followed by the word
acid (ex. H₂SO₃ -- X=sulfite = sulfurous
acid) CHAPTER 9 Chemical Names And
Formulas Test Review ...Writing the Names
and Formulas for Bases. A base is a
compound that produces OH⁻ in water.*

When naming a base, you name it like any other compound that starts with a regular or transition metal. Ex: NaOH = sodium hydroxide. When writing the formula for a base, remember to balance charges. Ex: magnesium hydroxide = Mg(OH)₂

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SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268-270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268-269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds. Name Date Class CHEMICAL NAMES AND FORMULAS 9 SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the

following binary ionic compounds. sodium sulfide g. CuCl₂ h. SnCl₄ i. MnO₂ c. d. SrBr₂

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hydroxide, OH⁻ 46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and CHAPTER 9 Study Guide - Quia Chapter 9 Chemical Names Formulas Chapter 9 - Chemical Names and Formulas. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. bibiloni1901. Terms in this set (20) Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence Chapter 9 Chemical Names Formulas Quiz Answer Key Chapter 9 "Chemical Names and Formulas" Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Help; Use these activities to help yourself to learn the vocabulary and major topics presented in this chapter. A B; anion: any atom or group of atoms that has a negative charge: Quia - Chapter 9 "Chemical Names and Formulas" Chemistry (12th Edition) answers to Chapter 9 - Chemical Names and Formulas - 9 Assessment - Page 299 80 including work step by step written by community members like you. Textbook Authors: Wilbraham, ISBN-10: 0132525763,

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CHAPTER 9 Assessment 42. a. 2- b. 1+ c. 1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d.1+ 44. a. barium ion b. iodide ion c. silver ion d. mercury(II) ion 45. cyanide, CN⁻ and hydroxide, OH⁻ 46. a. hydroxide ion b.

lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and
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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the anion name ends in -ite the acid name is the stem of the anion with the suffix -ous followed by the word acid (ex. H₂SO₃ -- X=sulfite = sulfurous acid)
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First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula.

Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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