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Chem Worksheet A simple calorimeter
constructed from Styrofoam coffee cups,
such as you will use in the laboratory,
measures reaction heats under constant
pressure conditions; thus, $q_{rxn} = \Delta H_{rxn}$,
the change in enthalpy of the reaction.
This is often used to measure the heat
change of a solution formed in the inner
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much heat is produced when 85 g of sulfur
reacts according to the reaction below?
 $2S + 3O_2 \rightarrow 2SO_3$ $H = -792 \text{ kJ}$. - the H value
given in the equation is the amount of
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is the experimental measurement of heat
(q) produced in chemical and physical
processes. Heat can not be measured
directly, but temperature changes can be
measured. The factor that links these two
is heat capacity. Heat capacity, C, is
defined as the heat required to raise the
temperature of a University of Illinois at
Urbana-Champaign Dr.
Gupta/Thermochemistry/Practice/Calorime
try and Heats of Reaction/Page 3 of 3 7)
Use the equations given to calculate the
enthalpy change for the equation given
below. $2NO_2(g) \rightleftharpoons N_2O_4(g)$ $H = ?$ (Ans:
-24.0 KJ) Given: a) $N_2(g) + 2O_2(g) \rightleftharpoons N_2O_4(g)$ $H = +9.2 \text{ KJ}$ b) $N_2(g) + 2O_2(g) \rightleftharpoons 2NO_2(g)$ $H = +33.2$
KJ Thermochemistry/Practice-Calorimetry
and Heat of Reaction ... Name:
Thermochemistry Worksheet #1 1. The
reaction of magnesium with sulfuric acid
was carried out in a calorimeter. This
reaction caused the temperature of 27.0
grams of liquid water, within the
calorimeter, to raise from 25.0 C to 76.0 C.
Calculate the energy associated with this
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0 Period .. 0.903 . 0.449 4.18 ass . Lead
San . 0.386 0.128 0.740 / 4.18 J/oc *
Mtn 70 Hz e . 1. Three different 30-gram
metal samples brass, and #2 were heated
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— Calorimetry Calorimetry is the
experimental measurement of heat (q)
produced in chemical and physical
processes. Heat can not be measured
directly, but temperature changes can be
measured. The factor that links these two
is heat capacity. Heat capacity, C, Enthalpy
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Capacity, and Specific Heat. The heat
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required to raise the temperature, ΔT , of
an object by 1 o C. The three variables are
related by the equation $[q = C\Delta T]$ The
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try and Heats of Reaction/Page 3 of 3 7)
Use the equations given to calculate the
enthalpy change for the equation given
below. $2\text{NO}_2(\text{g}) \rightarrow \text{N}_2\text{O}_4(\text{g})$ $H = ?$ (Ans:
-24.0 KJ) Given: a) $\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{N}_2\text{O}_4(\text{g})$ $H = +9.2$ KJ b) $\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2(\text{g})$ $H = +33.2$ KJ

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Molar Heat Capacity, and Specific Heat.
The heat capacity, C , is the amount of
heat, q , required to raise the temperature,
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variables are related by the equation $[q = C\Delta T]$ The value of C in this ...

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(q) produced in chemical and physical
processes. Heat can not be measured
directly, but temperature changes can be
measured. The factor that links these two
is heat capacity. Heat capacity, C , is
defined as the heat required to raise the
temperature of a

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links these two is heat capacity. Heat
capacity, C ,

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Heat Ca acities Joules/ o Period .. 0.903 .
0.449 4.18 ass . Lead San . 0.386 0.128
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and Zn were heated to
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A simple calorimeter constructed from Styrofoam coffee cups, such as you will use in the laboratory, measures reaction heats under constant pressure conditions; thus, $q_{rxn} = \Delta H_{rxn}$, the change in enthalpy of the reaction. This is often used

to measure the heat change of a solution formed in the inner cup.

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