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Introduction. The flame test is one of the most commonly used analytical processes in chemistry. It is widely used to detect and analyze the presence of certain elements in the given salt or compound. Primarily, the flame test detects the presence of metal ions in a compound, and as ions of each element have a specific characteristic based in their emission spectrum, the flame test for every element is different and distinctive.

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First, prepare your lab by placing the goggles over your eyes, connecting the bunsen burner to the gas, heating the bunsen burner with the lighter, and placing wooden sticks inside of the elements. Then, place one of the saturated sticks into the flame. Finally, observe the various colors that will appear based on the element that is tested.

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Background: A flame test is used to detect the presence of certain metal ions. The test involves heating a sample of the element and observing the resulting color of the flame. When atoms of elements are heated to high temperatures, some electrons may absorb enough energy to allow them to move to higher energy levels.

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This activity is called a flame test and it's a real procedure used in labs. Its purpose is to identify specific elements in a material.

When the boric acid was in the flame, you probably notice a bright green portion of the flame. You may have seen it only briefly but it was there.

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To perform flame tests of metal cations in order to observe their characteristic colors, To perform calculations to determine the frequency and energy of the emitted photons. To relate these results to the types of electronic transitions occurring in these elements.

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Every element has a unique flame test color. It is a traditional art of the chemistry laboratory to use these colors to identify specimens of compounds that contain unknown metals.

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