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Watson Lloyd Institute 2.05 watsong@tcd.ie Adsorption / Emission spectra for Hydrogen Johann Balmer (1885) measured line spectra for hydrogen 364.6 nm (uv), 410.2 nm (uv), 434.1 nm (violet), 486.1 nm (blue), and 656.3 nm (red).An introduction to Molecular Orbital Theory.pptAccording to MO theory, the first sigma orbital is lower in energy than either of the two isolated atomic 1s orbitals - thus this sigma orbital is referred to as a bonding molecular orbital. The second, sigma-star (σ^*) orbital is higher in energy than the two atomic 1 s orbitals, and is referred to as an antibonding molecular orbital (in MO theory, a star (*) sign always indicates an ...1.11: The Nature of Chemical Bonds- Molecular Orbital ...The molecular orbital (MO) theory is a powerful and extensive approach which describes electrons as delocalized moieties over adjacent atoms. 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Introduction to Molecular Orbital Theory

A molecular orbital (MO) is an allowed spatial distribution of electrons in a molecule that is associated with a particular orbital energy. Unlike an atomic orbital (AO), which is centered on a single atom, a molecular orbital extends over all the atoms in a molecule or ion. Hence the molecular orbital theory of bonding is a delocalized approach.

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The molecular orbital (MO) theory is a powerful and extensive approach which describes electrons as delocalized moieties over adjacent atoms. The applications of the MO theory extend beyond the limitations of the Valence Shell Electron Pair Repulsion (VSEPR) model and the Valence Bond theory.

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