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oxygen to produce
aluminum oxide as
follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
a. If you use 2.3
moles of Al, how many
moles of Al_2O_3 can you
make? b. If you want 3.9
moles of Al_2O_3 , how
many moles of O_2 are
needed? 2. Chemistry
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answer key. a. Two

formula units KClO_3
decom- pose to form two
formula units KCl and
three molecules O_2 . b
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test b answer key. Four
molecules NH_3 react with
six molecules NO to form
five mol-Chapter 12
Stoichiometry Test B
Answer Key This chart
shows you that your score
doesn't plummet with
every question you can't
answer confidently
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chapter test b answer
key. You can do very well
on this test without
knowing or answering

everything. The key to
doing well on the SAT II
Chemistry is to follow a
strategy that ensures you
will see and answer all the
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12 g 7. How many grams
of nitric acid, HNO_3 , can
be prepared from the
reaction of 138 g of NO_2
with 54.0 g H_2O ...
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2) 19. The mass of
element X found in 1.00
mole of each of four
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with 54.0 g H_2O ... Practice Test Ch3 Stoichiometry (page 2 of 2) 19. The mass of element X found in 1.00 mole of each of four
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formula units KClO_3
decom- pose to form two

formula units KCl and
three molecules O_2 . b
Chapter 12 stoichiometry
test b answer key. Four
molecules NH_3 react with
six molecules NO to form
five mol-

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12 Test A Answers**

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ensures you will see and
answer all the .

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produce aluminum oxide
as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
a. If you use 2.3
moles of Al_2O_3 , how many
moles of Al_2O_3 can you
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moles of Al_2O_3 , how
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