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Quantity (Common Name/s) (Common) Symbol/s SI Units

Dimension Number of molecules N: dimensionless dimensionless

Number of moles n: mol ...  $C V = \partial / \partial J K -1 [M][L ...$

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 systems: [°R]= 9 5 [K] [°R]= [°F]+459.67 Of course, one can then  
 write that [°F]= 9 5 [°C]+32 1.2.3 PRESSURE Pressure, , is also an  
 intensive property.

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 equilibrium (VLE) describes the distribution of a chemical species  
 between the vapor phase and a liquid phase.. The concentration  
 of a vapor in contact with its liquid, especially at equilibrium, is  
 often expressed in terms of vapor pressure, which will be a partial  
 pressure (a part of the total gas pressure) if any other gas(es ...

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